

Significant Anti-Inflammatory Properties of a Copper(II) Fenoprofenate Complex Compared with its Parent Drug. Physical and Chemical Characterization of the Complex

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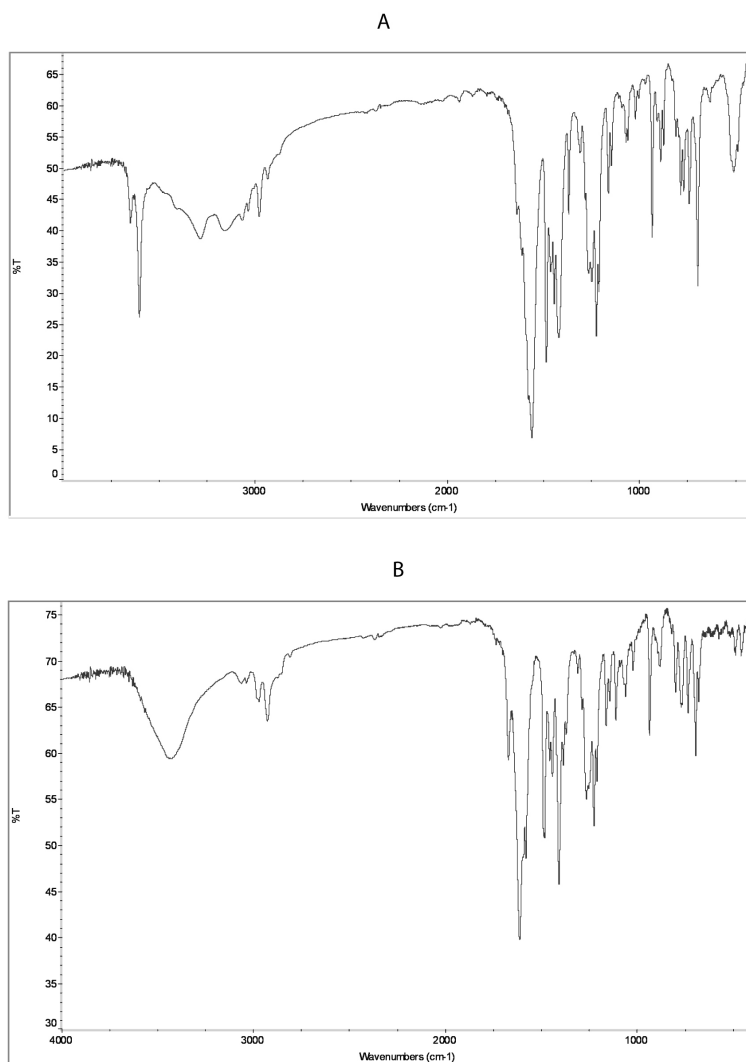


Figure S1. IR spectra of KBr dispersions of $[\text{Ca}(\text{fen})_2]\cdot\text{H}_2\text{O}$ (A) and microcrystalline $[\text{Cu}_2(\text{fen})_4(\text{dmf})_2]\cdot 2\text{H}_2\text{O}$ (B).

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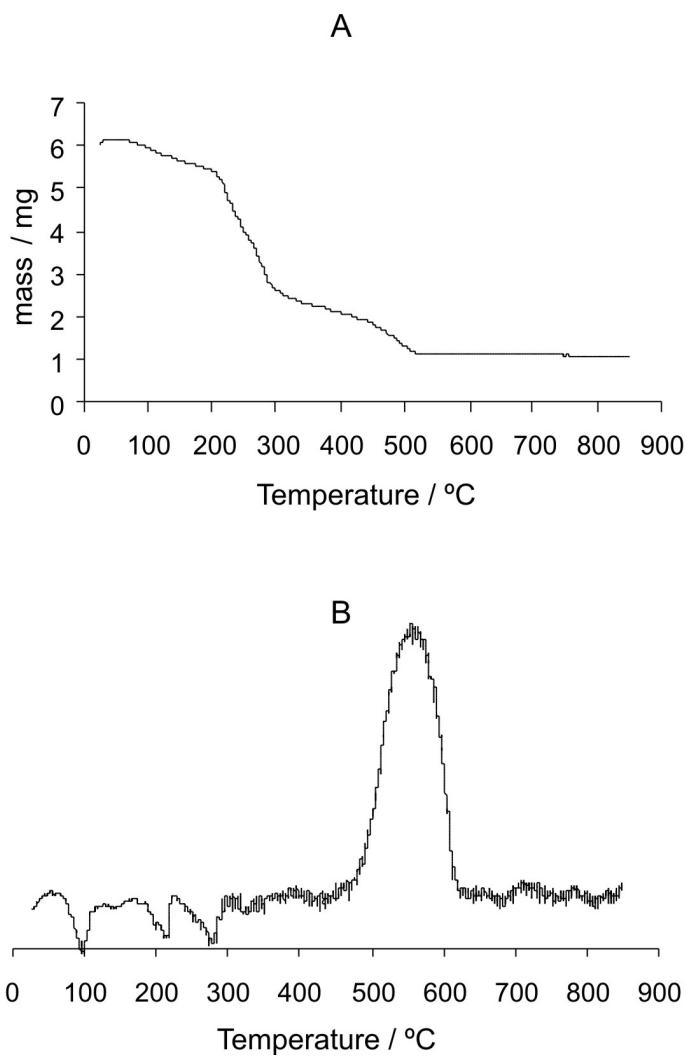


Figure S2. TGA (A) and DTA (B) curves recorded for single crystals of $[\text{Cu}_2(\text{fen})_4(\text{dmf})_2]$ (6.0 mg) under N_2 flux ($20 \text{ cm}^3 \text{ min}^{-1}$). Heating rate: $5 \text{ }^\circ\text{C min}^{-1}$.

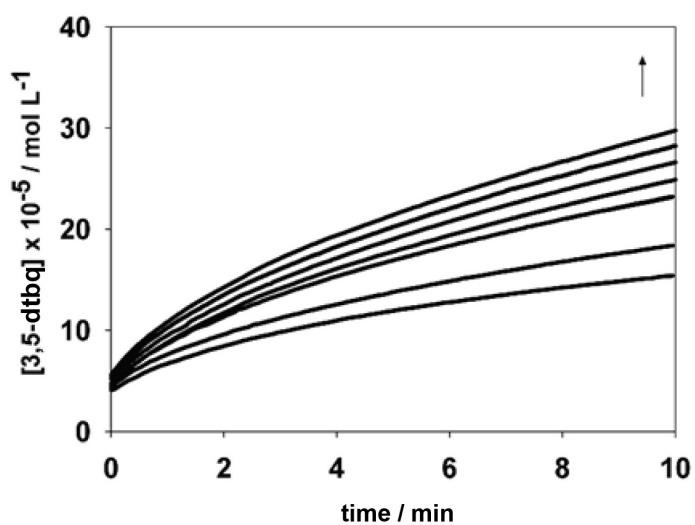


Figure S3. Variation of the 3,5-dtbq concentration vs. time. The up arrow indicates increasing concentrations of the substrate (0.125; 0.25; 0.50; 1.0; 2.0; 3.0; 4.0; 5.0 mol L^{-1}). Copper(II) complex concentration: $2.50 \times 10^{-5} \text{ mol L}^{-1}$. Temperature: $25.0 \pm 0.5 \text{ }^\circ\text{C}$.

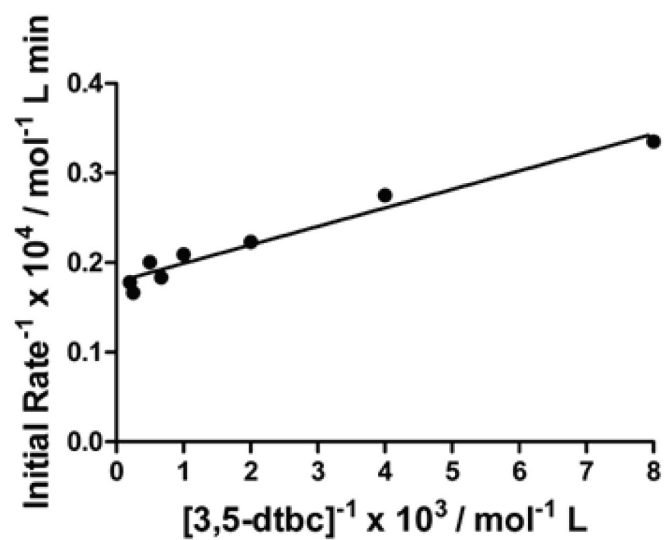


Figure S4. Lineweaver-Burk correlation plot for the catecholase mimetic activity of the copper(II) fenpropfenate complex in methanolic solution.

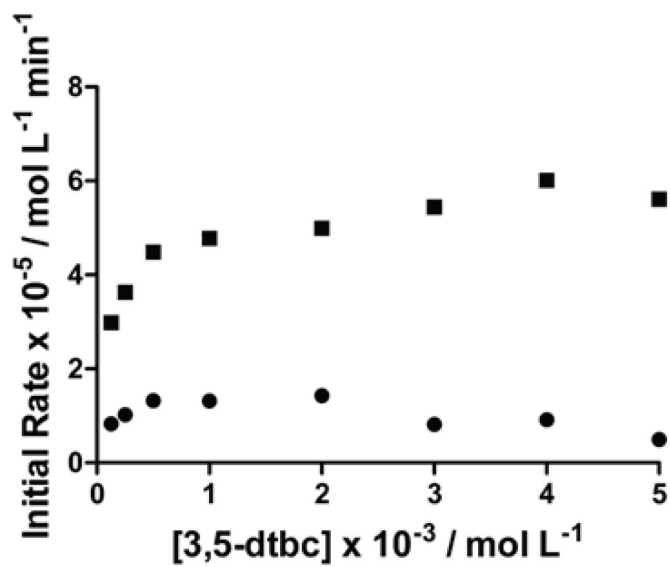


Figure S5. Comparative catecholase mimetic activity for the copper(II) fenpropfenate complex (■) and free cupric ions (●) under the same experimental conditions.