

## Easy Epoxidation of Monoterpenes from Common Starting Materials

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Epoxidation of monoterpenes,  $\alpha$ -pinene,  $\beta$ -pinene, limonene,  $\alpha$ -terpinene, and (*R*)-carvone was carried out by the *in situ* production of a peroxyacid rather than direct addition of such an expensive and difficult to handle chemical. Previous reports showed use of metal catalysts with high yields, while methodologies without catalysts at high temperature showed yields lower than 30%. The authors report a methodology that produces peroxyacetic acid *in situ* yielding up to 75% pure epoxide at room temperature avoiding the use of catalysts. The products were analyzed by gas chromatography mass spectrometry (GC-MS), and structures were characterized by <sup>1</sup>H and <sup>13</sup>C nuclear magnetic resonance (NMR).

**Keywords:** epoxidation, pinene, limonene, terpinene, carvone peroxyacetic acid

### Introduction

Epoxidation of olefins is important for industrial processes, because epoxides are easily converted into diverse products or functional groups. They are present in natural sources, such as leaves, flowers and bark.<sup>1-3</sup> Epoxides are frequently used in the perfume industry as starting materials or intermediates of fragrances,<sup>1</sup> which are complex mixtures that are normally composed of secondary metabolites like terpenoid structures, for example limonene,  $\alpha$ - and  $\beta$ -pinene, among others.<sup>4</sup> Essential oils also find application as solvents, production of lacquer, fibers like rayon, cosmetics, cleaning products and in the food industry.<sup>5</sup> Epoxides from fragrances are especially important because they are easily converted into glycols, amino alcohols, and esters.<sup>6-9</sup>

The synthesis of epoxides requires addition of expensive, unstable and hard to handle oxidizing reagents, normally involving an excess of peroxyacetic acid (PAA), 3-chloroperbenzoic acid (mCPBA), hydroperoxides, and dimethyldioxirane (DMDO).<sup>10,11</sup> Hydrogen peroxide is a convenient oxygen donor, but it requires the use of a metal catalyst. For example, tungsten, manganese, iron, and chromium catalyze epoxidation of olefins in high yields.<sup>12,13</sup> Palladium and gold catalysts are used

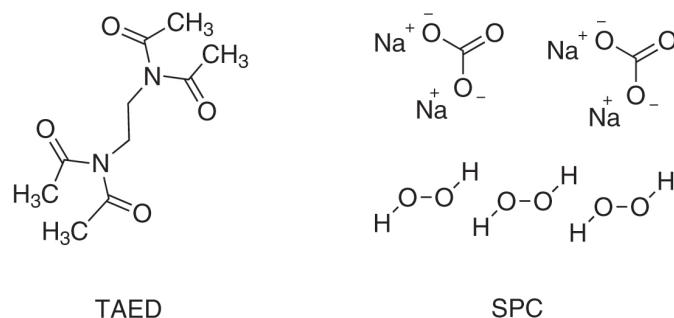
in combination with hydrogen peroxide for improving stereoselectivity.<sup>14</sup> Although homogeneous catalytic systems can become problematic during product isolation, heterogeneous catalysts such as oxide mixtures, zeolites,<sup>15-17</sup> and supported metals are good options for easy epoxide final purification.<sup>18</sup> The Juliá-Colonna epoxidation employs phase transfer catalysis for the asymmetric epoxidation of chalcones with high enantiomeric excess of the product.<sup>19</sup> Normally the addition of peroxyacetic acid yields about 30% of the respective epoxide.<sup>20</sup> We now report the easy and economical *in situ* production of the peroxyacid from tetraacetylenediamine (TAED) and sodium percarbonate (SPC) with yields up to 75% without using any metal catalyst at room temperature, for the epoxidation of monoterpenes.

Figure 1 shows the chemical structures of TAED and SPC, both of which are economical and easy to handle chemicals. SPC is commonly known as sodium peroxycarbonate or solid hydrogen, which is used as a bleach in laundry detergents, because it easily forms hydrogen peroxide once it reacts with water. TAED has been reported as a bleaching activator after reacting with SPC in an aqueous medium forming peroxyacetic acid.

### Experimental

Preliminary studies were performed with chemicals

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**Figure 1.** Chemical structures of TAED and SPC.

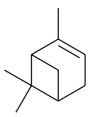
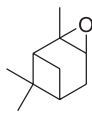
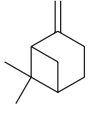
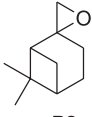


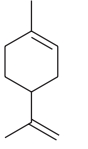
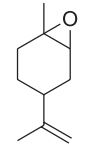
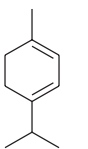
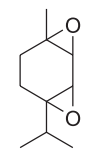
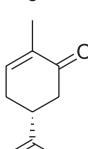
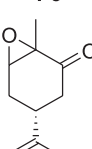
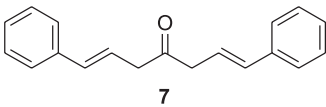
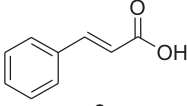
donated by Garcia and Millán Company (Cali, Colombia), which commercializes chemicals for bleaching and disinfection. Once the preliminary studies were successful, TAED, SPC and monoterpenes were purchased from Merck (Cali, Colombia), and they were used without any further purification. The following method for epoxidation of monoterpenes was optimized by changing parameters such as reaction time, solvent, temperature and molar ratio. Epoxidation of  $\alpha$ -pinene using the best experimental conditions is shown as described below, and this is the general method used for the epoxidation of all substrates, from **1** to **8** in Table 1.

#### Synthesis of the epoxide from $\alpha$ -pinene with TAED and SPC

In a 50 mL round bottom flask with magnetic stirring bar that contained 136.2 mg (1.0 mmol, 1.00 equivalent) of  $\alpha$ -pinene and 12.5 mL of dichloromethane (DCM), it was added 252 mg (3.0 mmol, 3.0 equivalents) of sodium bicarbonate, 414 mg (3 mmol, 3 equivalents) of SPC and 342.4 mg (1.5 mmol, 1.5 equivalents) of TAED. The round bottom flask was capped and left under vigorous stirring at room temperature. The progress of the reaction was monitored by thin layer chromatography (TLC) from the first hour. After 3 h of reaction, the organic and aqueous layers were separated, and extraction of the aqueous layer was carried out three times with 10 mL of DCM. The combined organic layers were washed with 40 mL of water and 40 mL of brine. The obtained organic layer was dried with sodium sulfate obtaining a crude product. The mobile phase used for TLC was ethyl acetate and petroleum ether 15% v/v and the chromatography plate was developed with an ethanolic solution of phosphomolybdic acid.

Purification by flash chromatography eluting with 5% ethyl acetate/petroleum ether yielded in all cases more than 75% pure epoxide. The following spectral properties were observed:  $^1\text{H NMR}$  (400 MHz,  $\text{CDCl}_3$ )  $\delta$  0.93 (s, 3H,  $\text{CH}_3$ ), 1.28 (s, 3H,  $\text{CH}_3$ ), 1.33 (s, 3H,  $\text{CH}_3$ ), 1.62 (d, 2H,  $J$  9.5 Hz,  $\text{CH}_2$ ), 1.68-1.74 (m, 1H, CH), 1.85-2.03 (m, 5H,  $\text{CH}_2$ ,  $\text{CH}_2$ , CH), 3.04 (d, 1H,  $J$  3.9 Hz, CH);  $^{13}\text{C NMR}$  (100 MHz,

**Table 1.** Monoterpenes and their respective epoxides and reaction yields

Substrate	Product	Yield / %
 <b>1</b>	 <b>P1</b>	75
 <b>2</b>	 <b>P2</b>	73
 <b>3</b>	 <b>P3</b>	68
 <b>4</b>	 <b>P4</b>	72
 <b>5</b>	 <b>P5</b>	57
 <b>6</b>	 <b>P6</b>	20
 <b>7</b>	N.P.	–
 <b>8</b>	N.P.	–

N.P.: no epoxidation product was observed.

CDCl<sub>3</sub>)  $\delta$  60.20, 56.82, 45.13, 40.50, 39.78, 27.63, 26.70, 25.84, 22.37, 20.14; high-resolution mass spectrometry (HRMS) (electron ionization (EI))  $m/z$ , calculated for C<sub>10</sub>H<sub>16</sub>O [M]<sup>+</sup>: 152.1201, found: 152.1198. Elemental analysis calculated for C<sub>10</sub>H<sub>16</sub>O: C, 78.90; H, 10.59; O, 10.51; found: C, 78.91; H, 10.60; O, 10.51.

Gas chromatography coupled to a mass detector (GC-MS)

A Shimadzu GCMS-QP-2010 was employed for separation of products and starting materials. The column was a Restek XTI-5. The following heating ramps were carried out: (i) initial temperature of 40 °C for 1 min, (ii) heating at 5 °C min<sup>-1</sup> up to 200 °C, (iii) heating at 10 °C min<sup>-1</sup> up to 300 °C. The mass spectrometer transfer line was maintained at 350 °C, and the potential of the ionization source (EI) was 70 eV. Mass spectra were repeatedly scanned from  $m/z$  30-250. The acquired mass spectrum was matched with the mass spectrum in the Shimadzu GCMS-QP-2010 local library.

Analysis by nuclear magnetic resonance (NMR)

NMR analyses of pure products were carried out in a Bruker instrument at 400 MHz for <sup>1</sup>H and 100 MHz for <sup>13</sup>C. All <sup>1</sup>H and <sup>13</sup>C NMR spectra were recorded in CDCl<sub>3</sub>. Chemical shifts were referenced to residual signals from CDCl<sub>3</sub> (7.27 ppm for <sup>1</sup>H, 77.23 ppm for <sup>13</sup>C) with 0.03% v/v tetramethylsilane (TMS) as an internal reference. The NMR splitting patterns are designated as follows: s, singlet; d, doublet; t, triplet; m, multiplet; br, broad signal.

## Results and Discussion

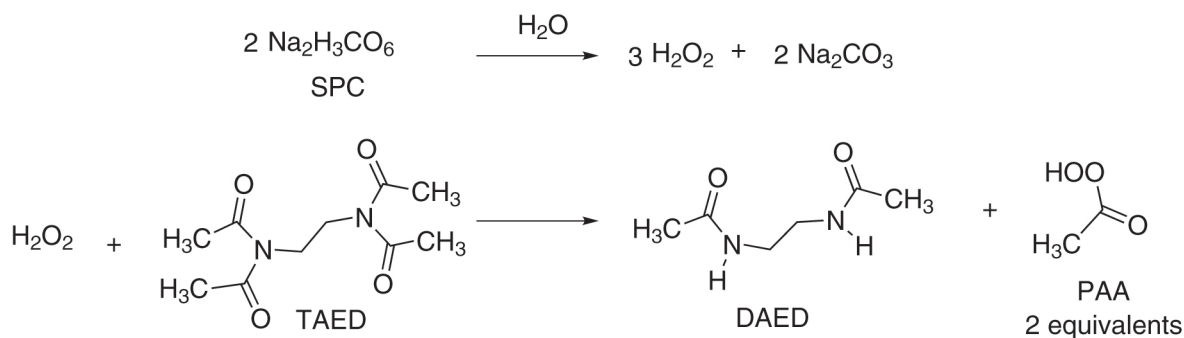
In order to test our proposed methodology, it was necessary to reproduce the reaction previously reported by Furniss *et al.*<sup>20</sup> in which  $\alpha$ -pinene was epoxidized by the addition of peroxyacetic acid (20% m/m) under constant stirring in a biphasic toluene/water solution at 40 °C during

3 h of reaction.<sup>21</sup> The reaction produced the expected epoxide, but even though an excess of peroxyacid was added during the course of the reaction, yields lower than 30% were obtained in all cases. In contrast, when the reaction was performed by *in situ* production of peroxyacetic acid from TAED and SPC in a mixture of water/DCM,<sup>22</sup> yields were improved dramatically, the epoxidation of monoterpenes yielded up to 75% pure epoxide. Although the following discussion is focused on the data obtained for the epoxidation of  $\alpha$ -pinene (**1**), information related to a series of monoterpenes and some model chemicals, from **2** to **8** in Table 1, showed similar results, which are available in the Supplementary Information (SI) section.

After one hour, the reaction progress for the epoxidation of  $\alpha$ -pinene was monitored by TLC, as the starting material began to disappear while the epoxide product, 2,7,7-trimethyl-3-oxatricyclo[4.1.1.0<sup>2,4</sup>]octane (**P1**), appeared.<sup>22</sup> Figure S1 (SI section) shows the TLC of the starting material and crude product. Although the mixture was reacted for periods longer than 3 h, no effect on the reaction yield was observed; no by-products were produced and only epoxide and  $\alpha$ -pinene were recovered.

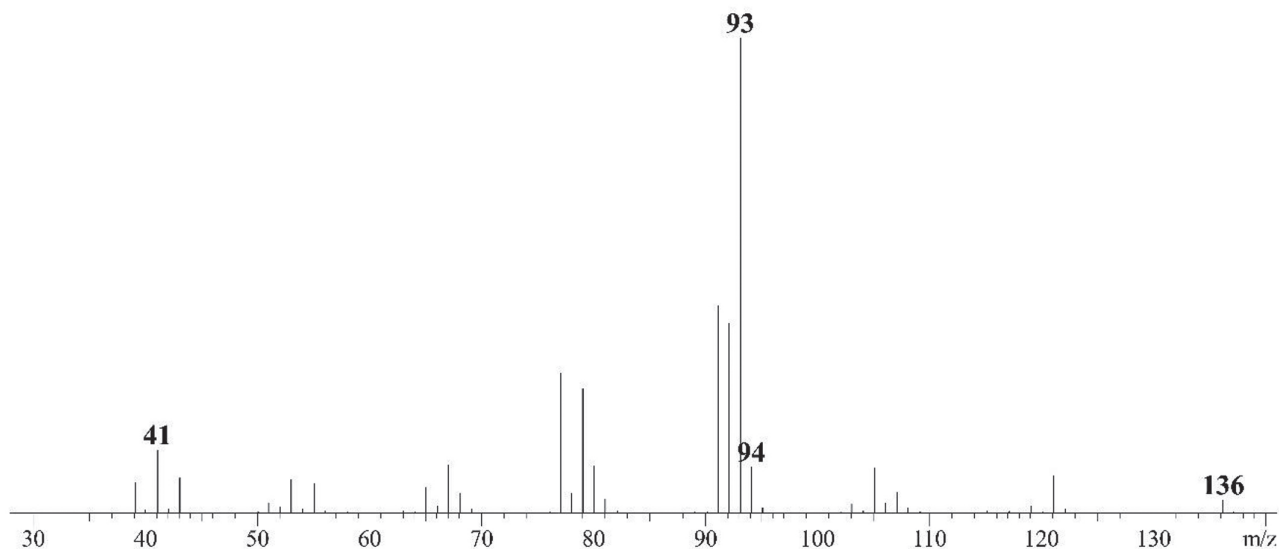
Scheme 1 shows the formation of peroxyacetic acid from TAED and SPC. While in previous reports PAA was added directly to the reaction, the *in situ* formation of PAA in the interphase of DCM and water for epoxidation has not been previously reported in the literature. In addition, it is important to note that the reaction produces two equivalents of PAA *per* every equivalent of TAED, improving the atom economy of the reaction.

Once the PAA is formed, it rapidly reacts with the unsaturation of the monoterpene, yielding the desired product epoxides (Scheme 2). Table 1 summarizes a series of monoterpenes, from **1** to **8**, and their respective epoxidized products, from **P1** to **P6**. It is important to highlight that cyclohexene (**3**), dibenzylideneacetone (**7**), and cinnamic acid (**8**) were also included in the series of monoterpenes as model compounds. In addition, limonene showed to be regioselective for the annular epoxidation as

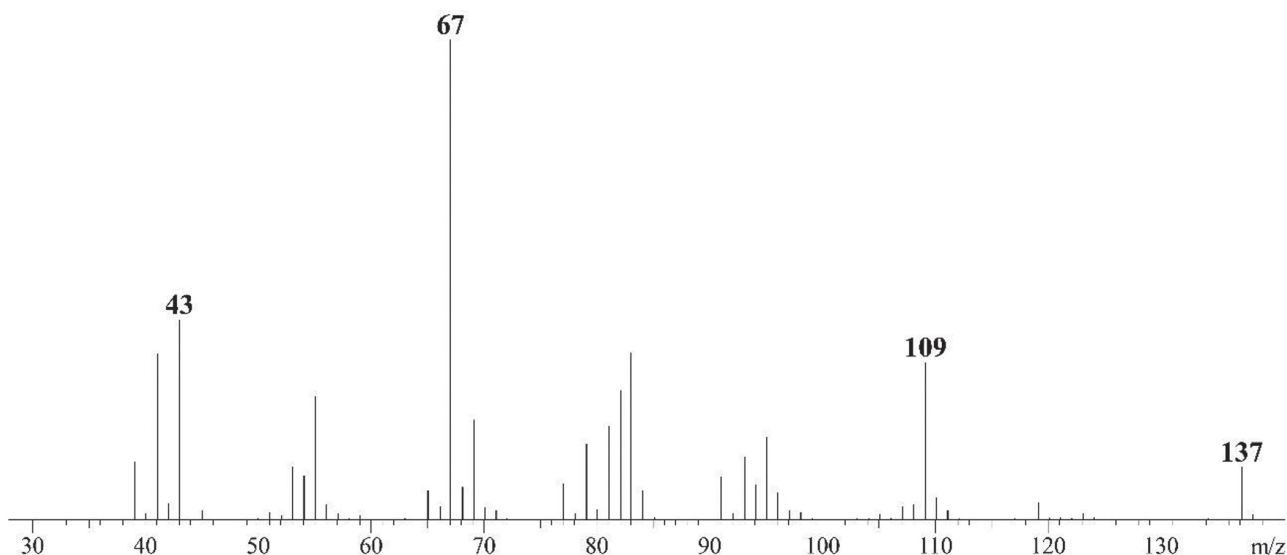


**Scheme 1.** Formation of peroxyacetic acid from SPC and TAED.





**Figure 3.** Mass spectrum of  $\alpha$ -pinene corresponding to the chromatographic peak at 6.49 min.



**Figure 4.** Mass spectrum of  $\alpha$ -pinene epoxide corresponding to the chromatographic peak at 9.35 min.

material. Figure 6 shows the  $^{13}\text{C}$  NMR spectrum of the epoxide product, which does not show any signal at low field corresponding to the double bond, 116 and 144 ppm, respectively, instead the carbons forming the oxirane ring show two signals at 57 and 60 ppm suggesting completion of the epoxidation (Figures 5 and 6 show portions of the spectra, while Figures S3 and S4 (SI section) show the full spectra with no traces of unsaturation).

All spectroscopic characterization of the product of  $\alpha$ -pinene epoxidation showed that the proposed reaction was successful, yielding 75% pure product. The same reaction conditions were applied to a series of monoterpenes as shown in Table 1. While  $\beta$ -pinene produced the respective epoxide in 73% yield, the model olefin, cyclohexene (**3**) produced the cyclohexene epoxide in 68% yield. In

addition to the monoterpenes with one double bond, the series of monoterpenes containing two olefin moieties, limonene (**4**),  $\alpha$ -terpinene (**5**), and (*R*)-carvone (**6**), were also exposed to the same reaction conditions for epoxidation. Limonene showed to be regioselective for annular epoxidation, yielding 72% of *cis* and *trans* product without the presence of diepoxidation. However, since  $\alpha$ -terpinene contains two olefins in the monoterpene ring, epoxidation of both olefins was observed, decreasing the yield to 57%, which was expected due to electronic and steric effects. Monoepoxidation was not observed, and a mixture of *cis* and *trans*  $\alpha$ -terpinene epoxides was obtained; the spectroscopic data are presented in the SI section. In order to understand the behavior of the inside and outside of the ring epoxidation for monoterpenes, (*R*)-carvone (**6**)

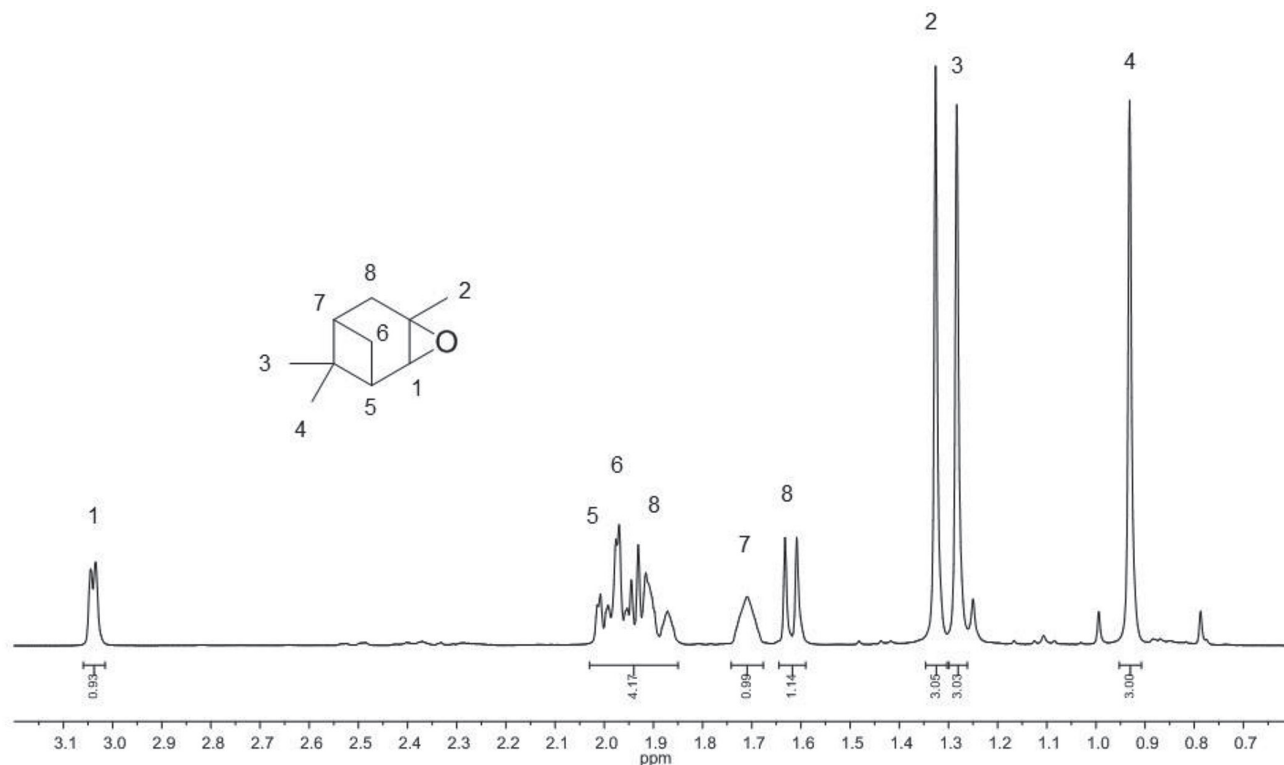


Figure 5. Expanded  $^1\text{H}$  NMR (400 MHz,  $\text{CDCl}_3$ ) of  $\alpha$ -pinene epoxide.

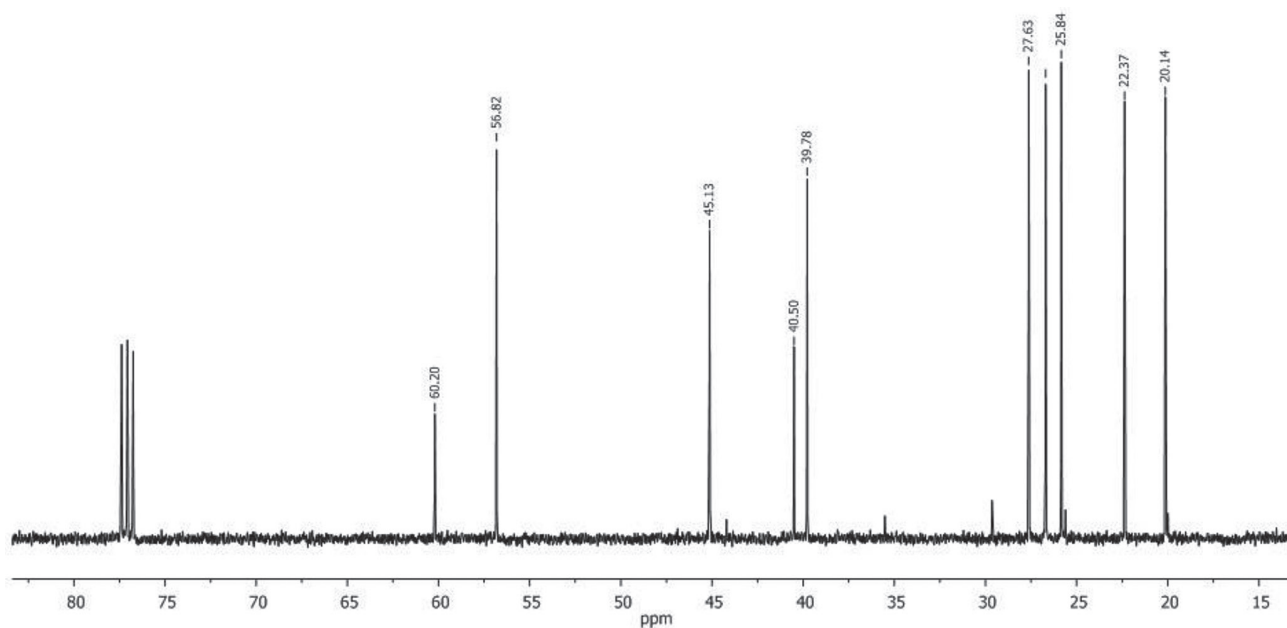


Figure 6. Expanded  $^{13}\text{C}$  NMR (100 MHz,  $\text{CDCl}_3$ ) of  $\alpha$ -pinene epoxide.

was studied. Epoxidation of **6** under the same reaction conditions yielded only 20% of product. Epoxidation of the olefin outside of the ring was not observed, which demonstrates that the reaction was indeed regioselective for the annular olefin. However, the epoxidation occurred at a low extent due to the presence of the carbonyl group, an electron withdrawing group that shifts electron density

away from the olefin. Spectroscopic characterization of the (*R*)-carvone epoxide is presented in the SI section. Since the presence of electron withdrawing groups affects the reaction, epoxidation of dibenzylideneacetone (**7**) and cinnamic acid (**8**) was performed. While in both cases epoxidation of the olefins was not observed, the cinnamic acid produced benzaldehyde and benzoic acid as undesired

products. In general, it was observed that olefins attached to electron donating groups yielded a larger extent of epoxide product, which was the case of olefins attached to alkyl groups, **1**, **2**, **4**, and their respective products, **P1**, **P2**, and **P4**. On the other hand, olefins attached to electron withdrawing groups, **6**, **7**, and **8**, led to formation of by-products or epoxides **P6** in low yields.<sup>24</sup>

## Conclusions

Successful synthesis of monoterpene epoxides by *in situ* production of peroxyacetic acid from sodium percarbonate and tetraacetylenediamine is reported, employing easy to access chemicals commonly used for disinfection processes. This methodology produces the desired epoxide without the use of catalysts in high yields, up to 75% depending of the monoterpene, without the need for excess peroxyacids, which are difficult to handle and expensive chemicals. The epoxidation reaction was found to be regioselective to the double bond in the ring, which was demonstrated by the annular epoxidation of limonene. The presence of electron withdrawing groups affected the formation of the desired epoxide, reducing the nucleophilic ability of the olefin to produce the oxirane ring. Additionally, this methodology is carried out at room temperature using a biphasic medium of dichloromethane and water that facilitates the final extraction of the product. These factors make this synthesis more environmentally friendly in comparison to reports using, for example: dimetyldioxirane,<sup>25</sup> peroxyacids,<sup>6-8</sup> tungstates<sup>26,27</sup> and alumina.<sup>24</sup> Other substrates are currently under study in order to expand the scope of these reaction conditions.

## Supplementary Information

Supplementary data are available free of charge at <http://jbcbs.s bq.org.br> as PDF file.

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